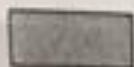




Atoms that lose electrons and have a positive charge



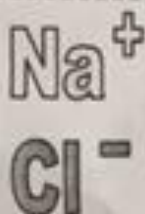
Atoms want to have 8 electrons in their outer shell = stable



Atoms that gain electrons and have a negative charge



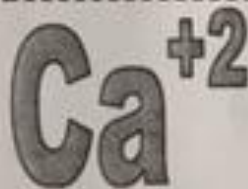
These are found to the left of the staircase and form ionic bonds with non-metals: give away electrons and become +  
Hydrogen is the exception.



Atoms with either a positive (+) or negative (-) electrical charge



These are found to the right of the staircase and form covalent bonds with each other, or take electrons away from metals to form ionic bonds and become -



The total number of electrons that an atom either gains or loses in order to form a chemical bond with another atom.



These touch the staircase in the periodic table and have properties of both metals and non-metals.



Type of bond where atoms give electrons to other atoms and there is an attraction between oppositely-charged ions



This is the type of bond where electrons are shared between atoms.  
 $\text{H}_2\text{O}$  – electrons are shared between H's & O.



The electrons found in the outermost shell of an atom. This is where the chemical reactions and bonding takes place.



The outermost shell of an atom